

Answers

Formula-Nomenclature Practice Grid for Ionic Compounds

Anion Cation	Fluoride	Chloride	Oxide	Sulfide	Sulfate	Sulfite	Carbonate	Chlorate
Sodium	NaF	NaCl	Na ₂ O	Na ₂ S	Na ₂ SO ₄	Na ₂ SO ₃	Na ₂ CO ₃	NaClO ₃
Potassium	KF	KCl	K ₂ O	K ₂ S	K ₂ SO ₄	K ₂ SO ₃	K ₂ CO ₃	KClO ₃
Magnesium	MgF ₂	MgCl ₂	MgO	MgS	MgSO ₄	MgSO ₃	MgCO ₃	Mg(ClO ₃) ₂
Calcium	CaF ₂	CaCl ₂	CaO	CaS	CaSO ₄	CaSO ₃	CaCO ₃	Ca(ClO ₃) ₂
Aluminum	AlF ₃	AlCl ₃	Al ₂ O ₃	Al ₂ S ₃	Al ₂ (SO ₄) ₃	Al ₂ (SO ₃) ₃	Al ₂ (CO ₃) ₃	Al(ClO ₃) ₃
Chromium(II)	CrF ₂	CrCl ₂	CrO	CrS	CrSO ₄	CrSO ₃	CrCO ₃	Cr(ClO ₃) ₂
Iron(II)	FeF ₂	FeCl ₂	FeO	FeS	FeSO ₄	FeSO ₃	FeCO ₃	Fe(ClO ₃) ₂
Iron(III)	FeF ₃	FeCl ₃	Fe ₂ O ₃	Fe ₂ S ₃	Fe ₂ (SO ₄) ₃	Fe ₂ (SO ₃) ₃	Fe ₂ (CO ₃) ₃	Fe(ClO ₃) ₃
Copper(I)	CuF	CuCl	Cu ₂ O	Cu ₂ S	Cu ₂ SO ₄	Cu ₂ SO ₃	Cu ₂ CO ₃	CuClO ₃
Copper(II)	CuF ₂	CuCl ₂	CuO	CuS	CuSO ₄	CuSO ₃	CuCO ₃	Cu(ClO ₃) ₂
Silver	AgF	AgCl	Ag ₂ O	Ag ₂ S	Ag ₂ SO ₄	Ag ₂ SO ₃	Ag ₂ CO ₃	AgClO ₃
Ammonium	NH ₄ F	NH ₄ Cl	(NH ₄) ₂ O	(NH ₄) ₂ S	(NH ₄) ₂ SO ₄	(NH ₄) ₂ SO ₃	(NH ₄) ₂ CO ₃	NH ₄ ClO ₃

Answers

Formula-Nomenclature Practice Grid for Molecular Compounds

	Fluorine		Chlorine		Oxygen		Sulfur	
	Formula	Name	Formula	Name	Formula	Name	Formula	Name
Boron(III)	BF ₃	boron trifluoride	BCl ₃	boron trichloride	B ₂ O ₃	diboron trioxide	B ₂ S ₃	diboron trisulfide
Silicon(IV)	SiF ₄	silicon tetrafluoride	SiCl ₄	silicon tetrachloride	SiO ₂	silicon dioxide	SiS ₂	silicon disulfide
Phosphorus(V)	PF ₅	phosphorus pentafluoride	PCl ₅	phosphorus pentachloride	P ₂ O ₅	diphosphorus pentoxide	P ₂ S ₅	diphosphorus pentasulfide
Tungsten(VI)	WF ₆	tungsten hexafluoride	WCl ₆	tungsten hexachloride	WO ₃	tungsten trioxide	WS ₃	tungsten trisulfide

Some TRANSITION METALS

can form molecular

compounds. These are

an exception to the rule

For our purposes, consider all T.M.'s

to form ionic compounds